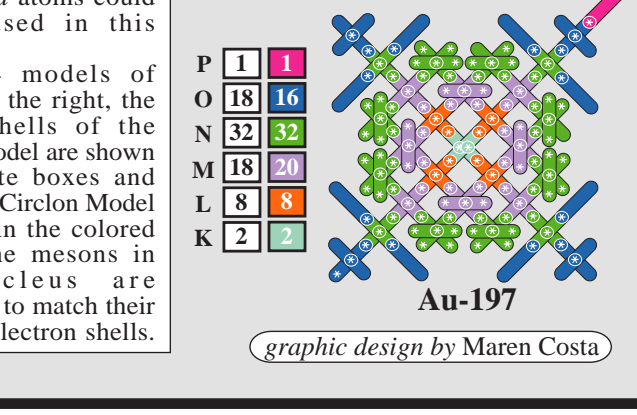
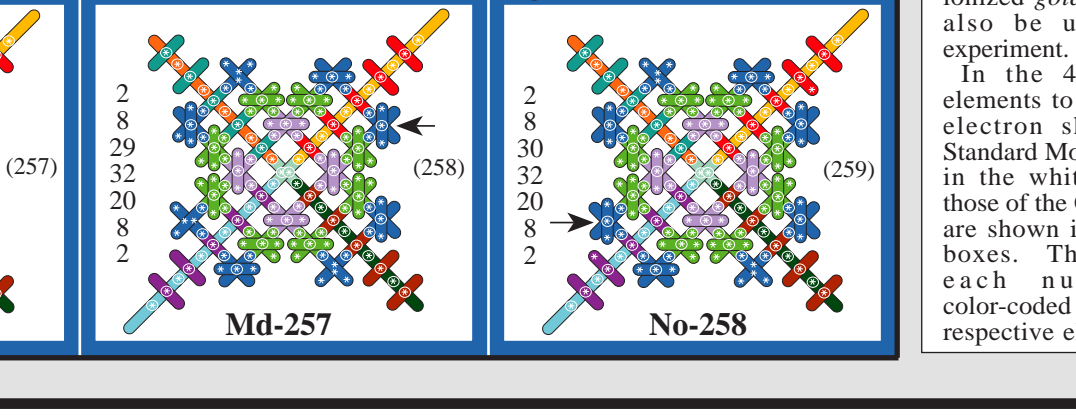
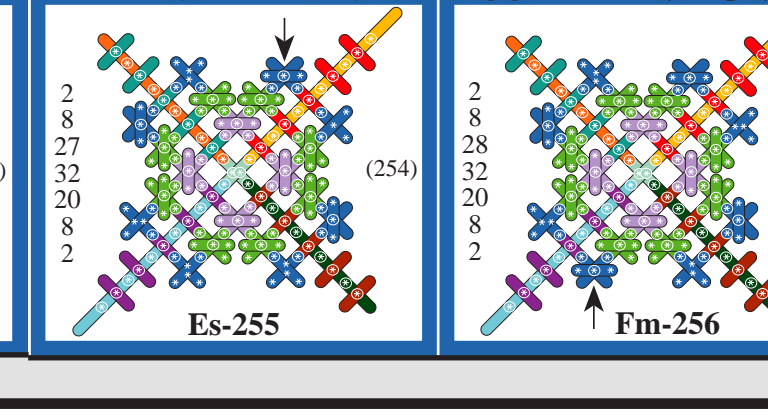
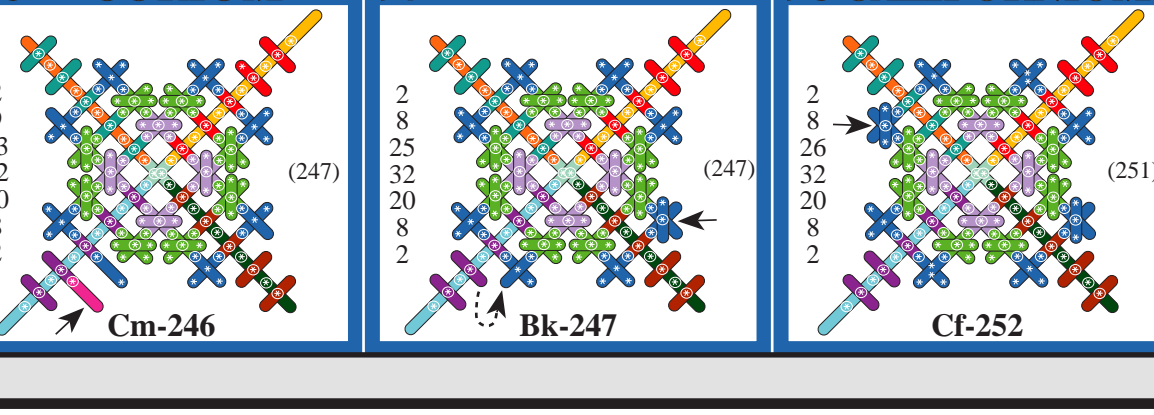
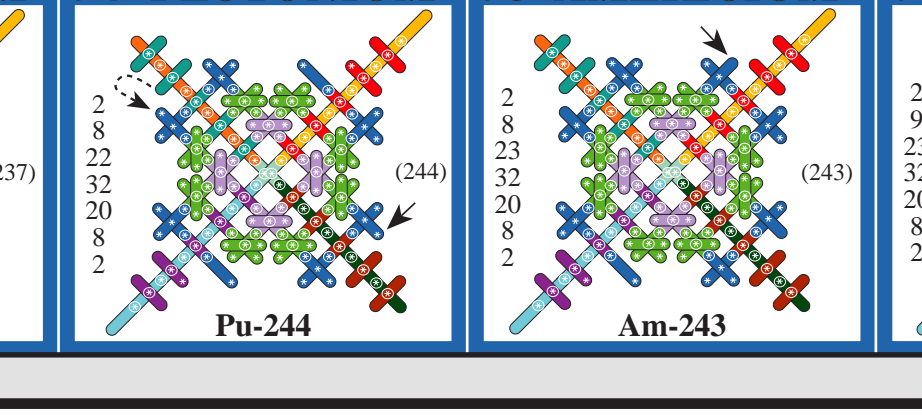
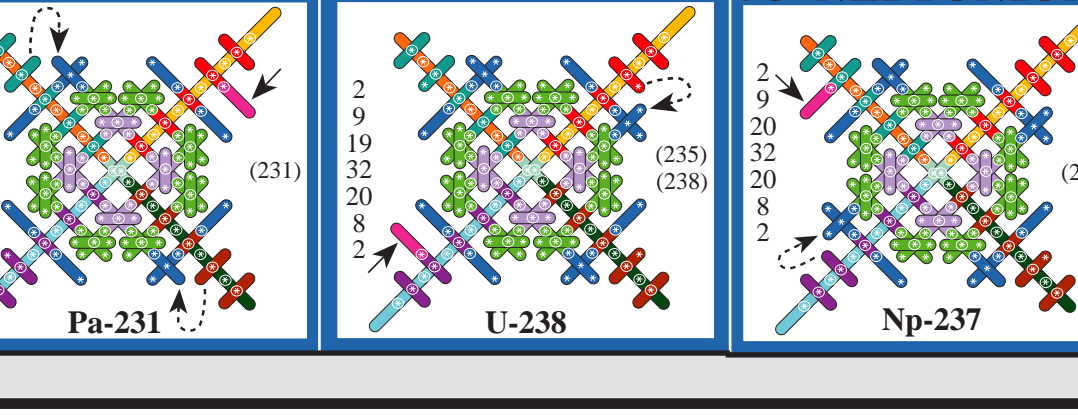
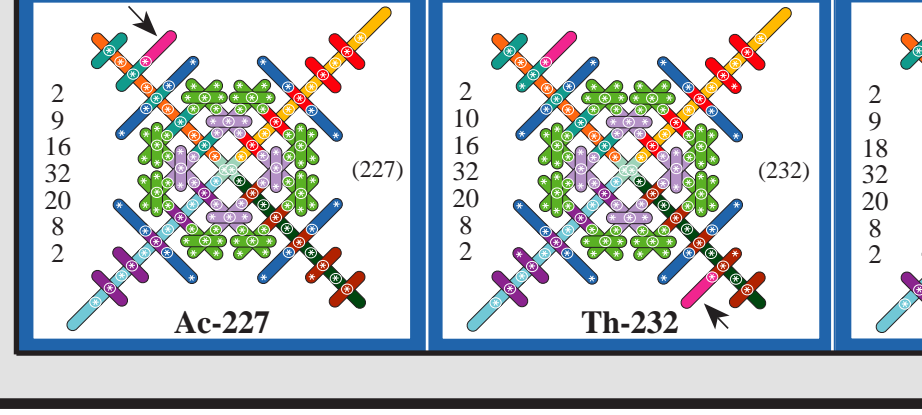
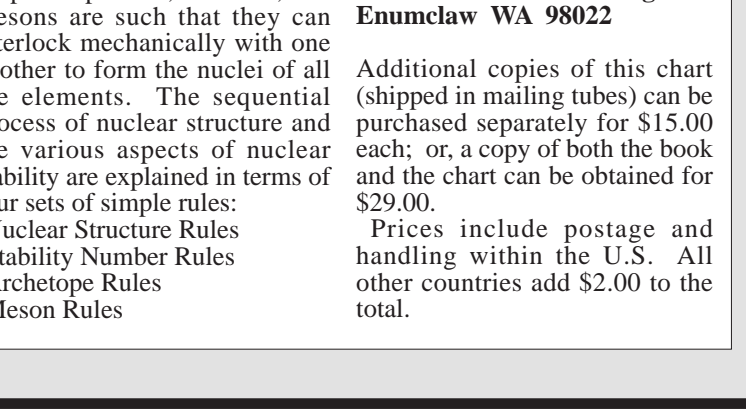
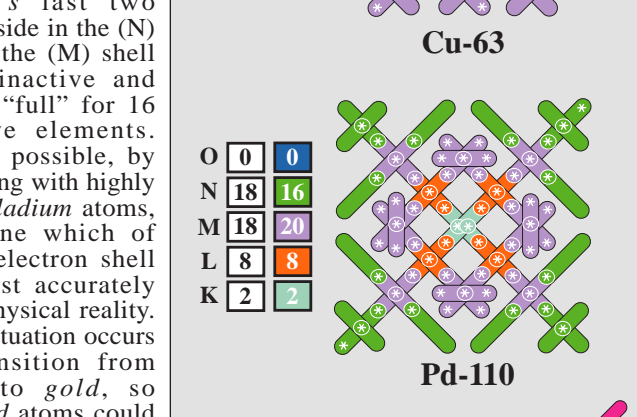
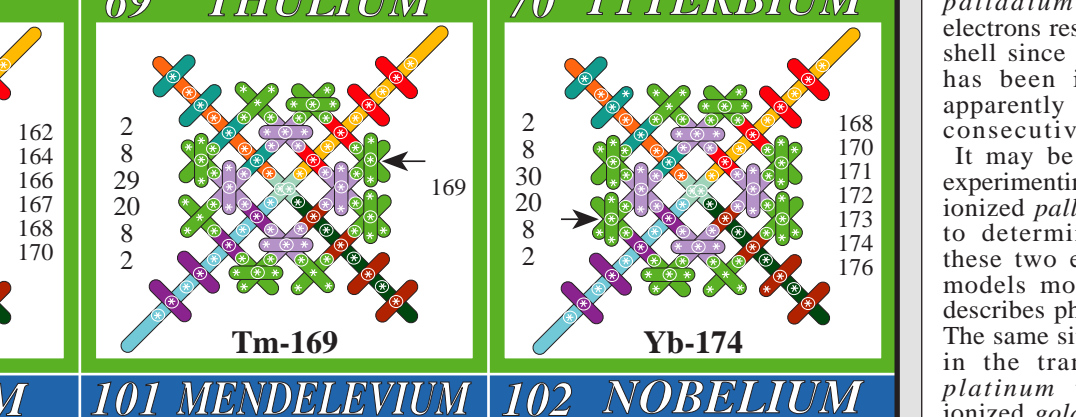
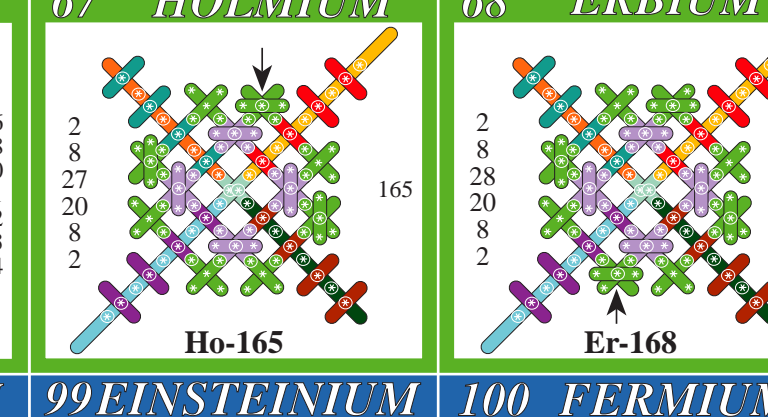
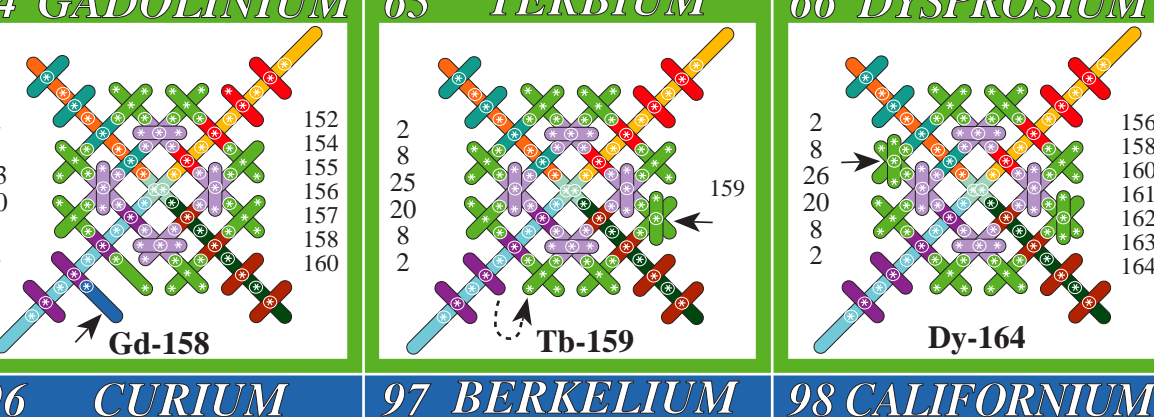
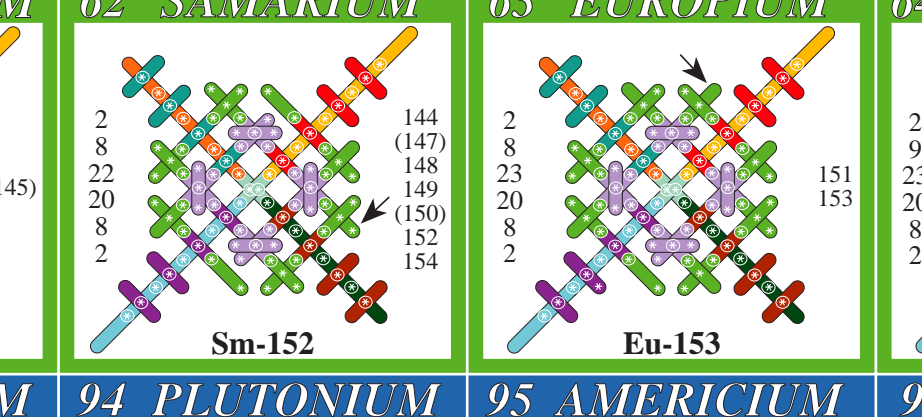
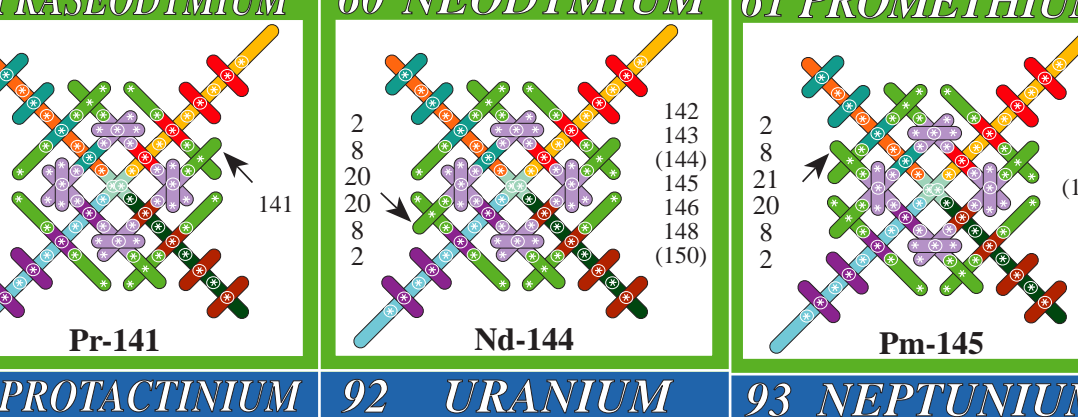
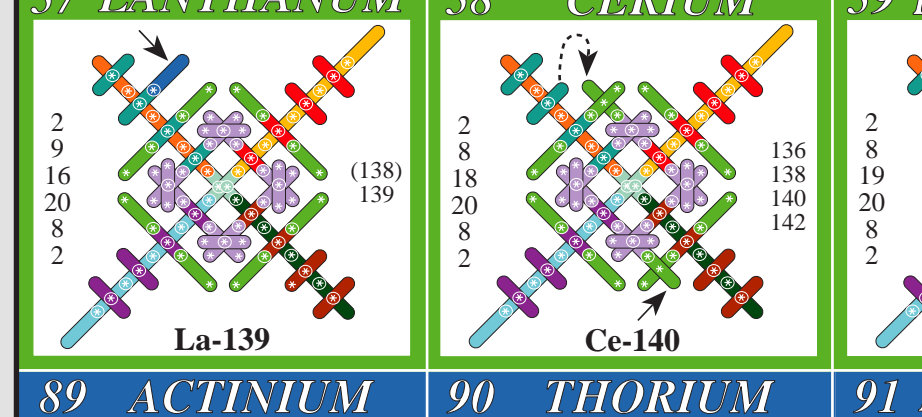
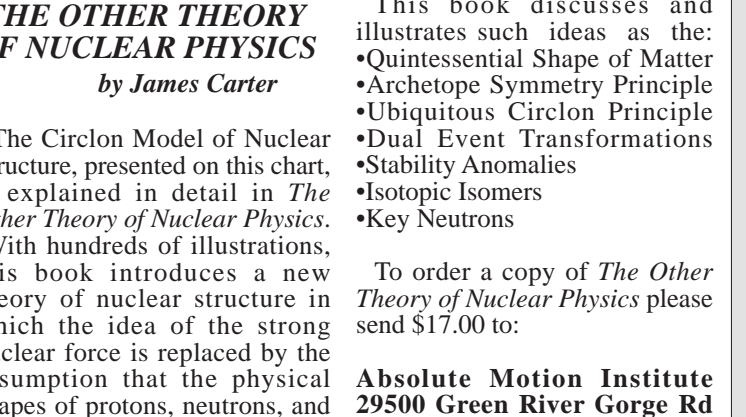
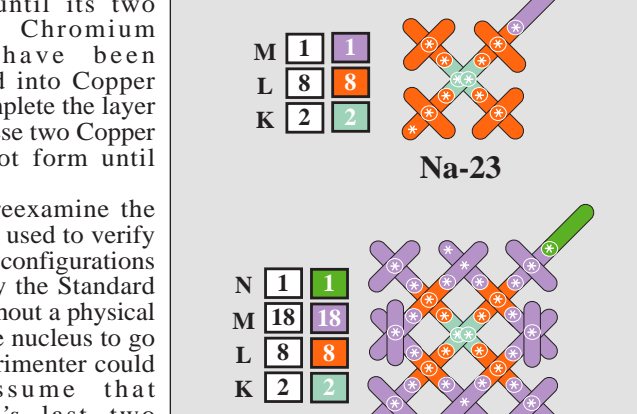
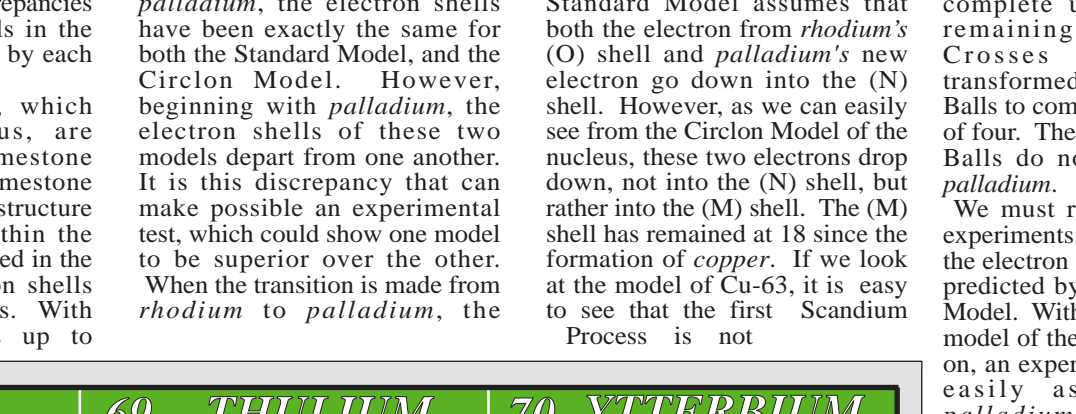
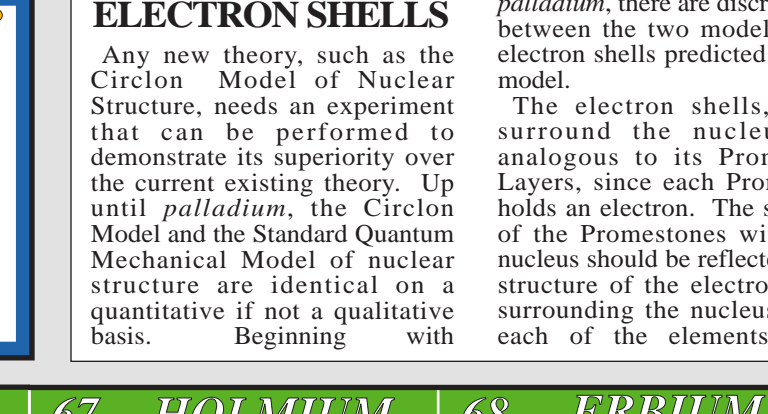
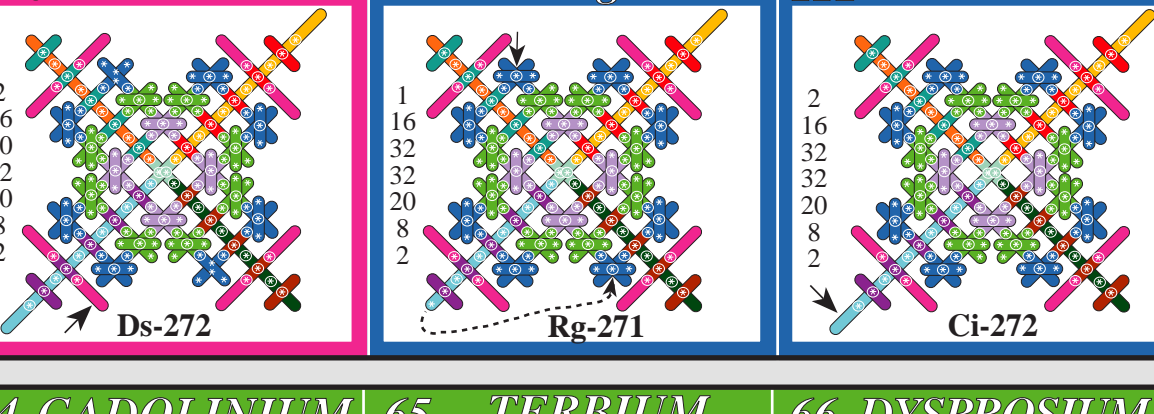
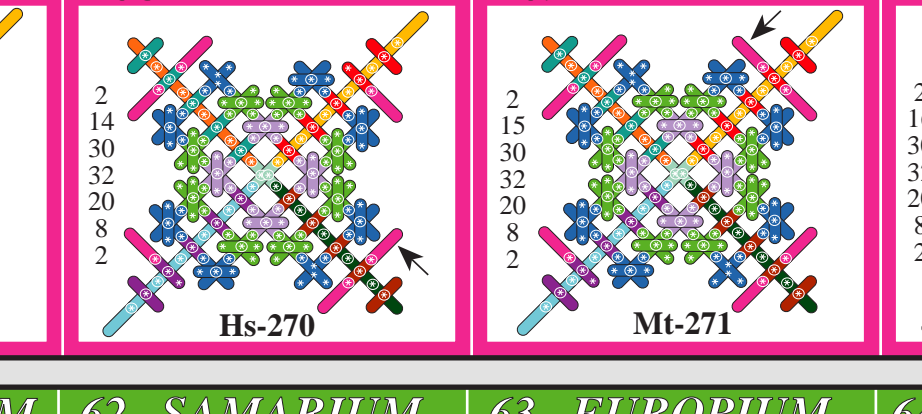
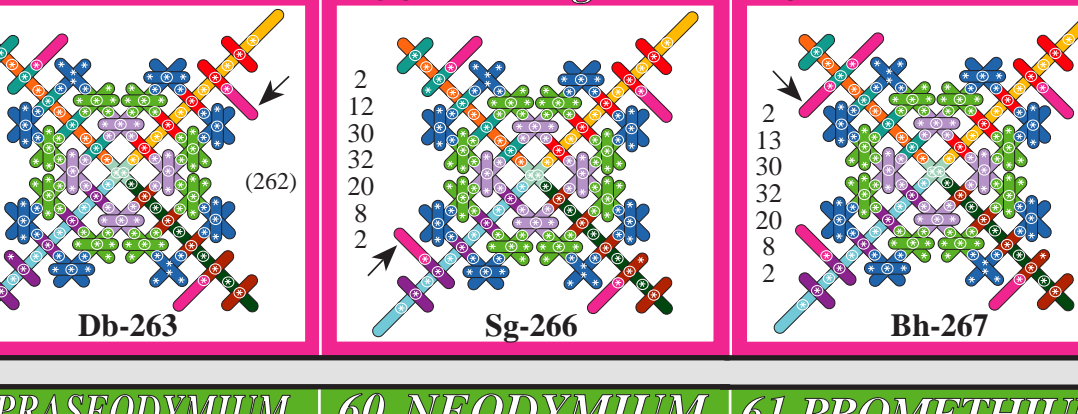
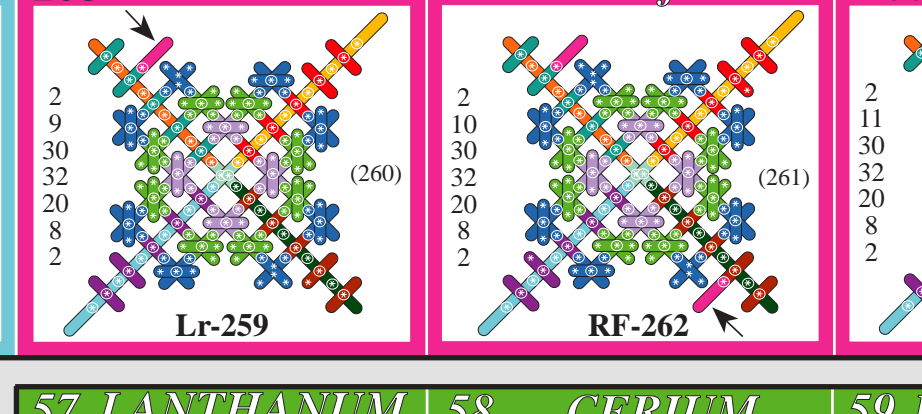
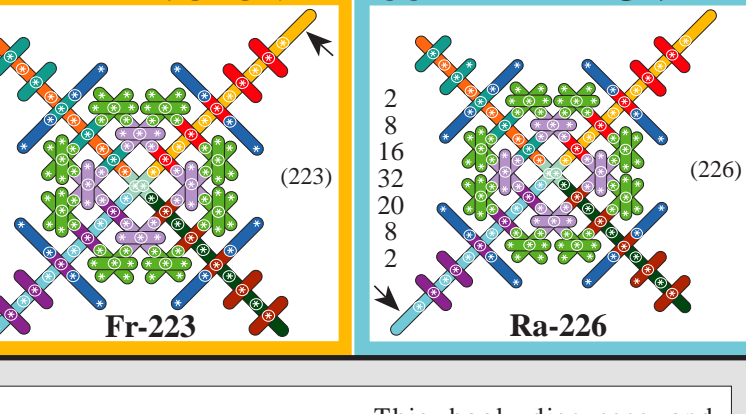
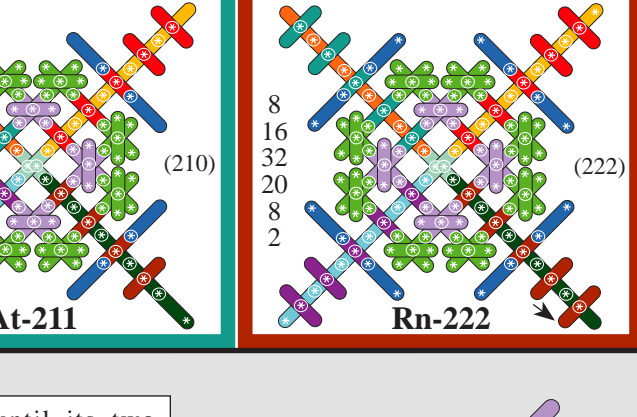
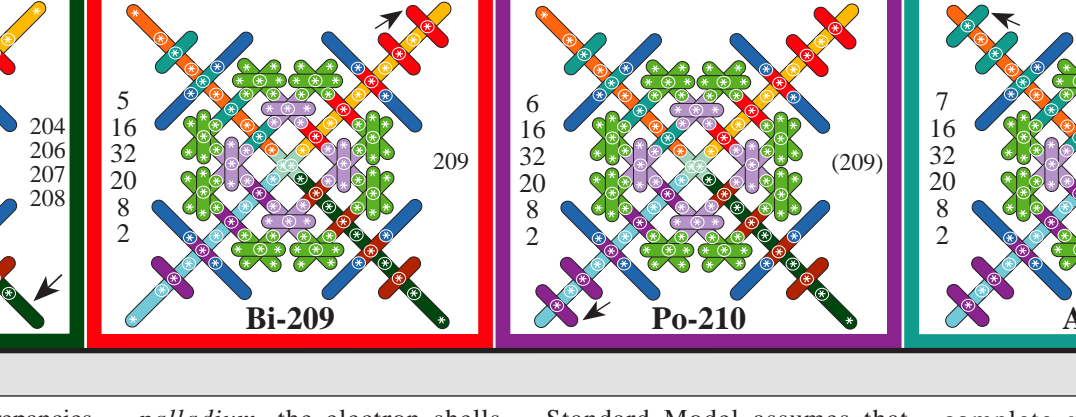
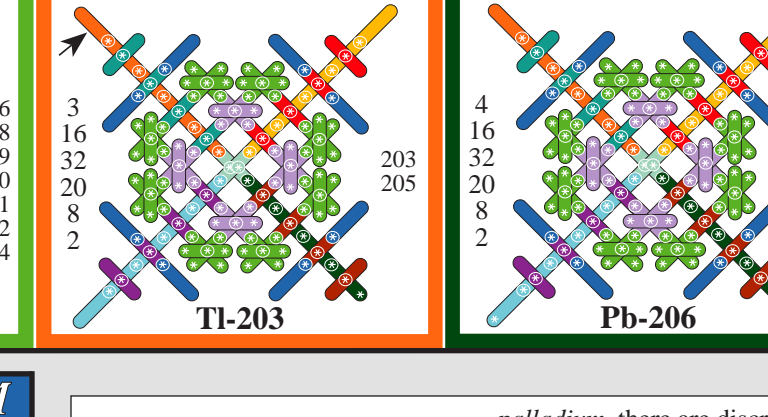
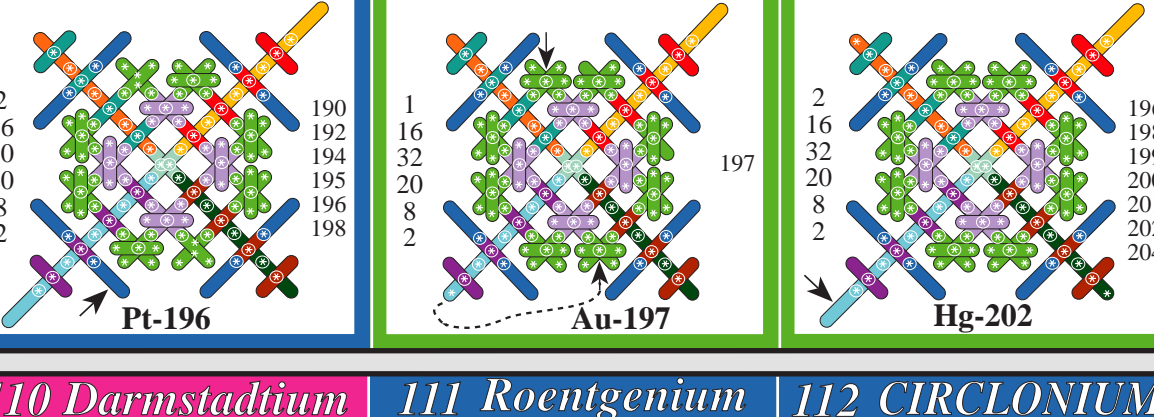
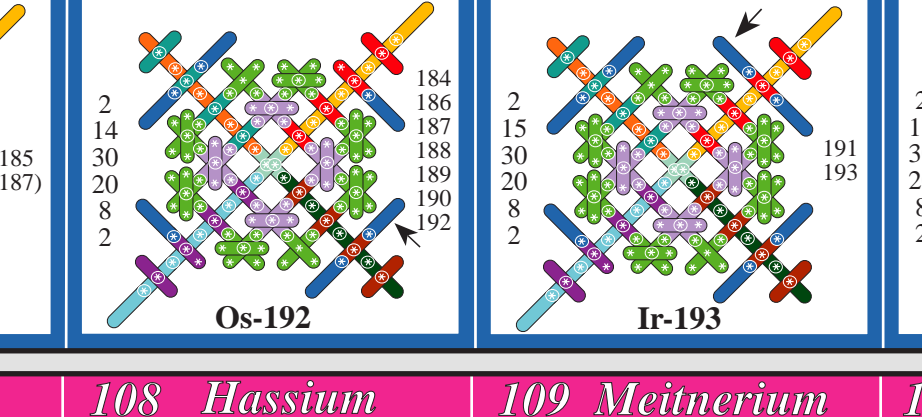
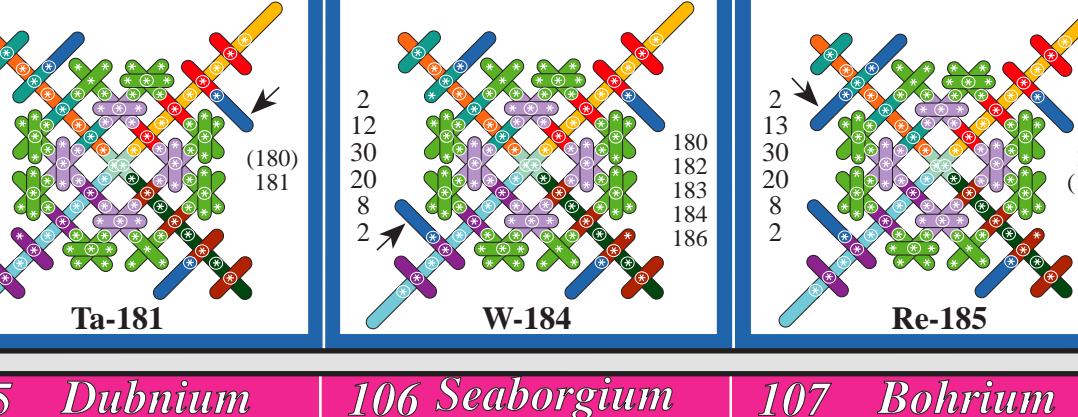
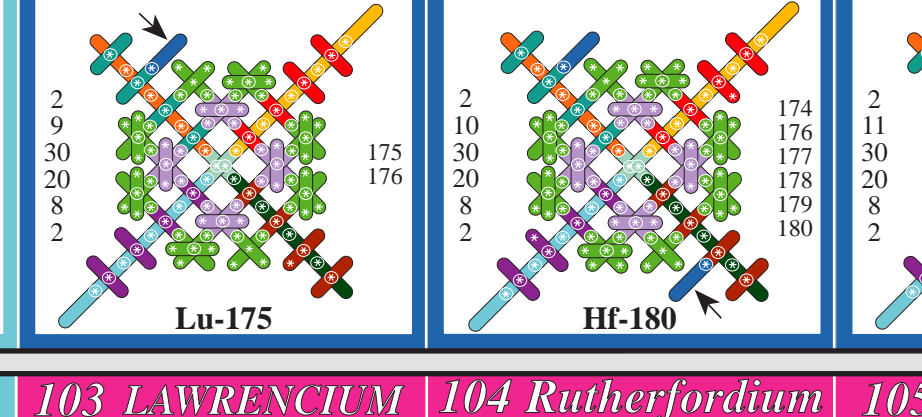
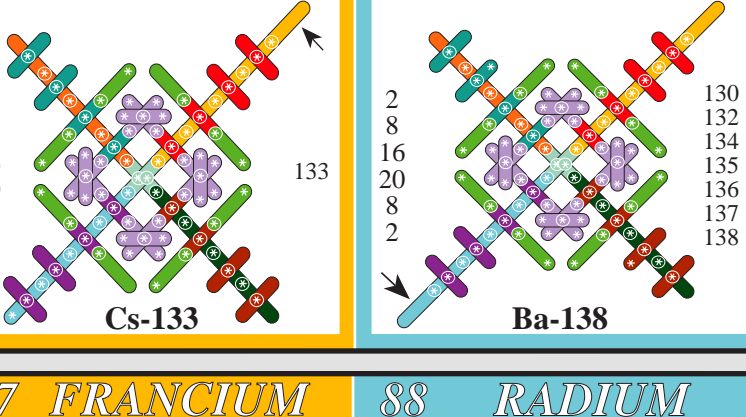
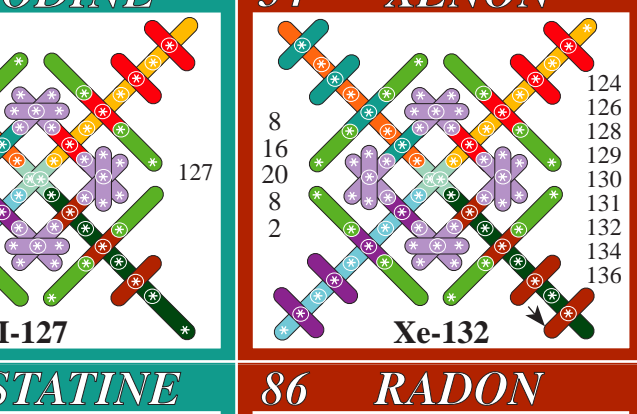
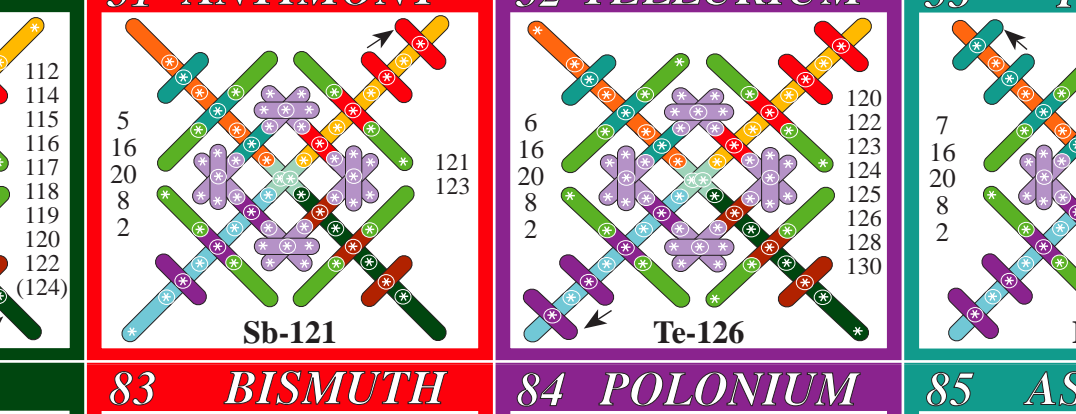
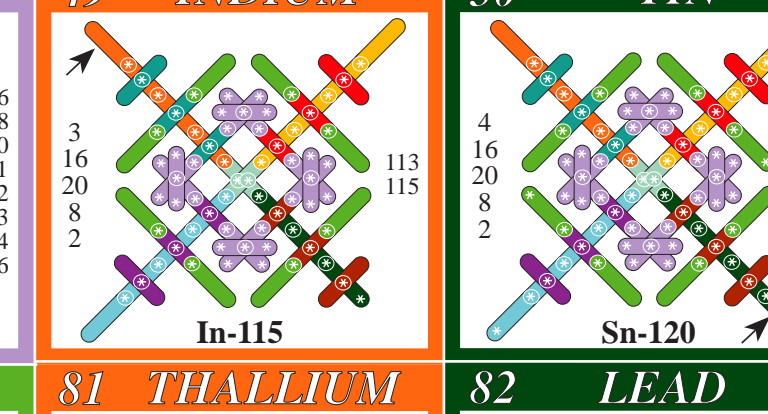
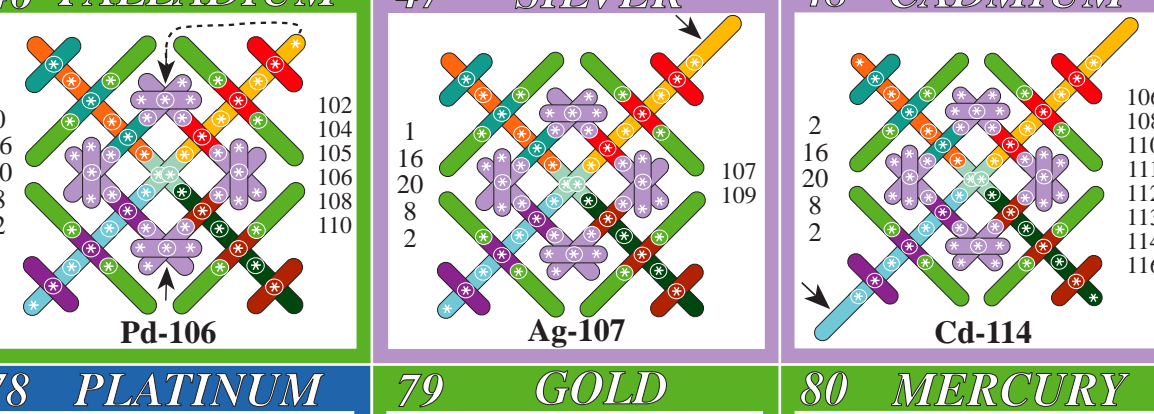
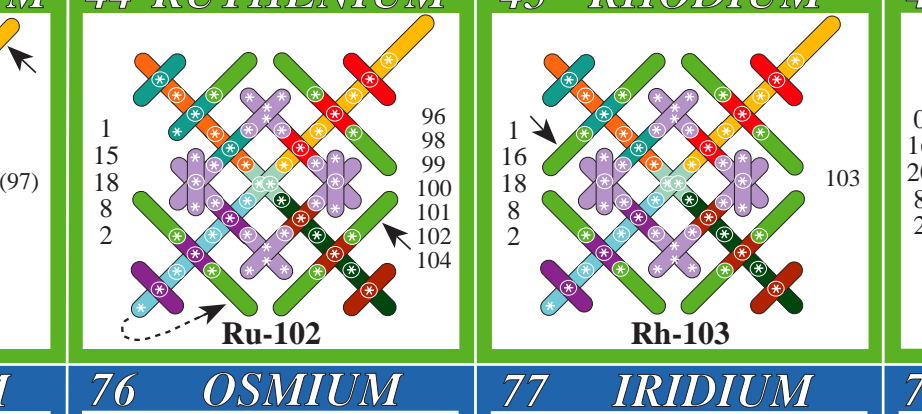
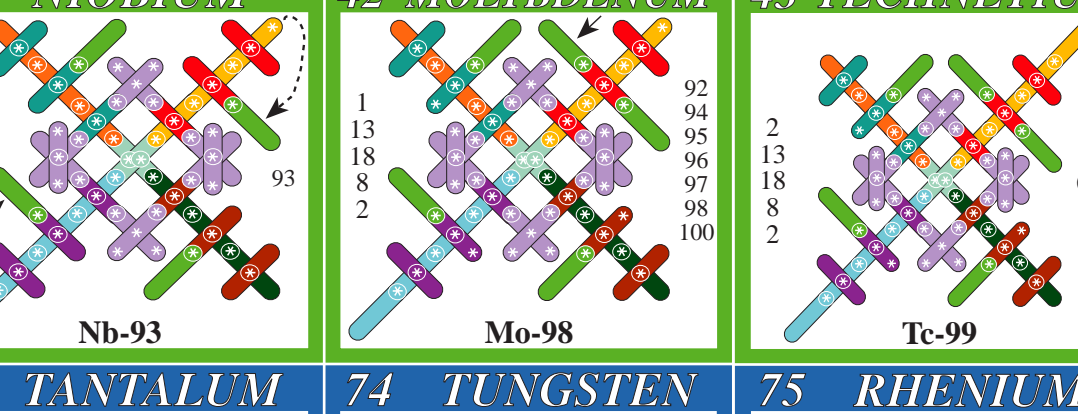
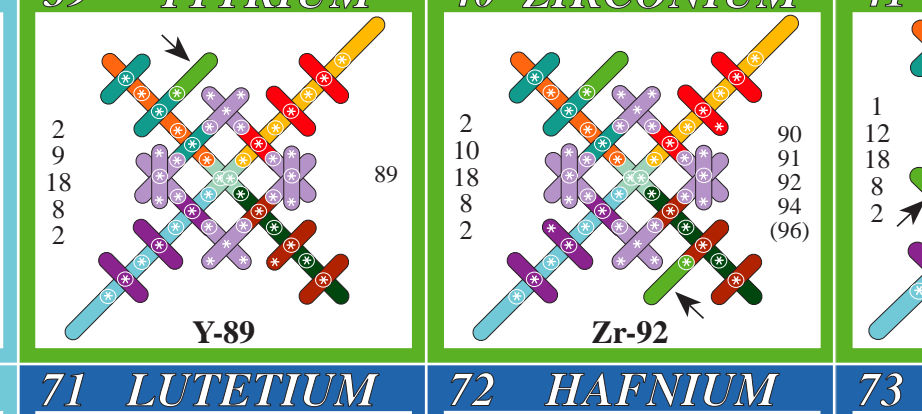
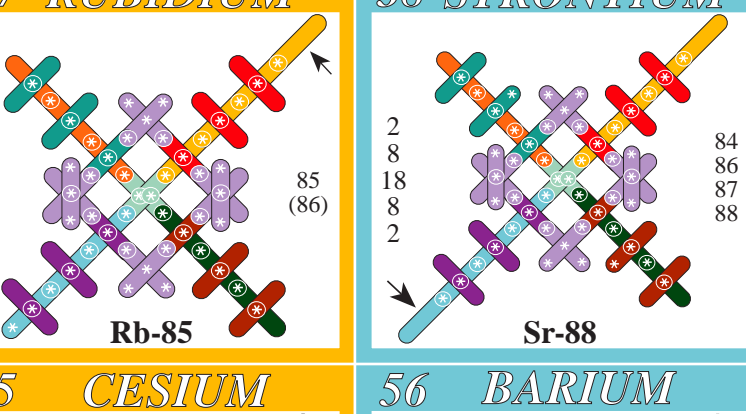
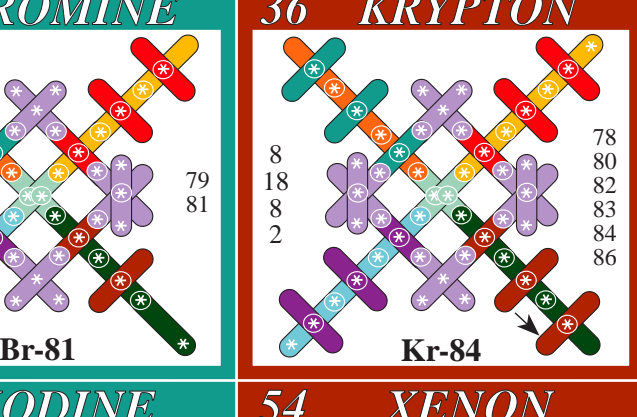
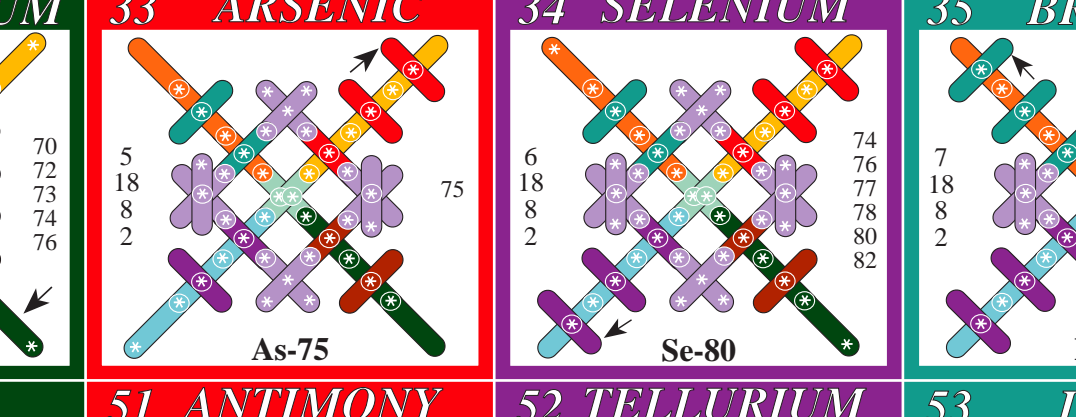
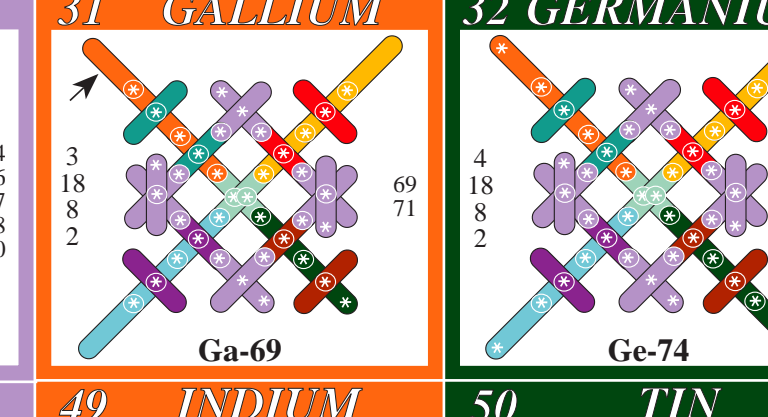
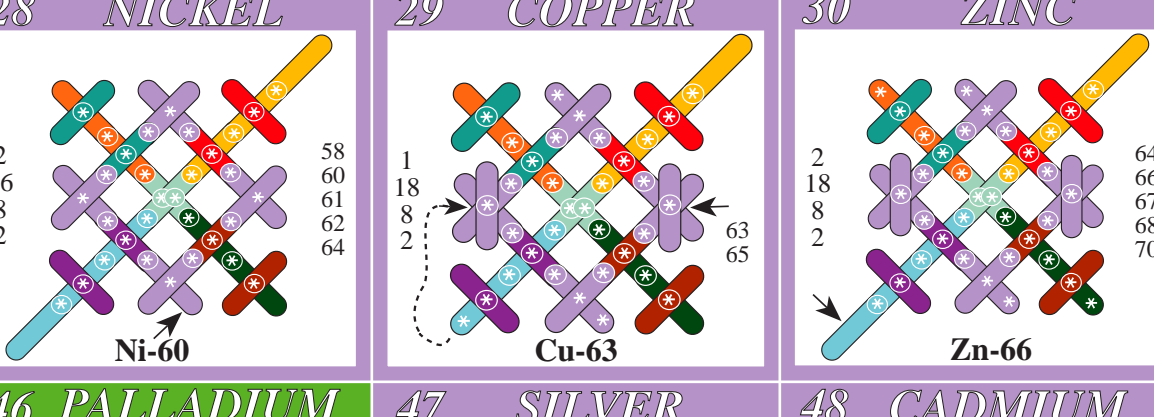
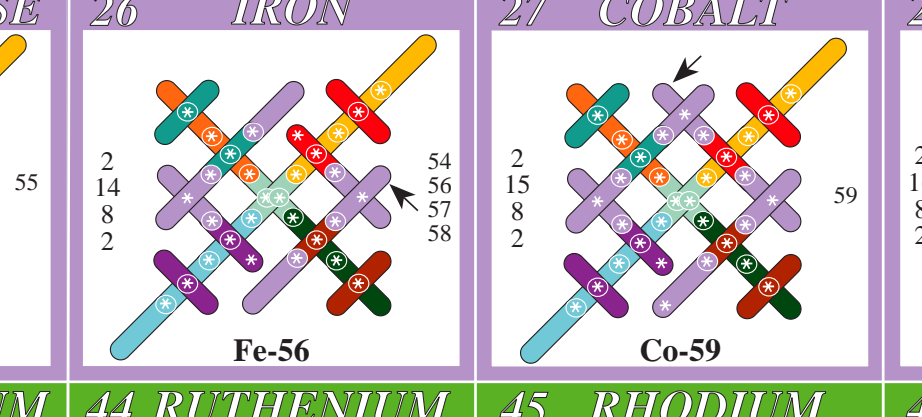
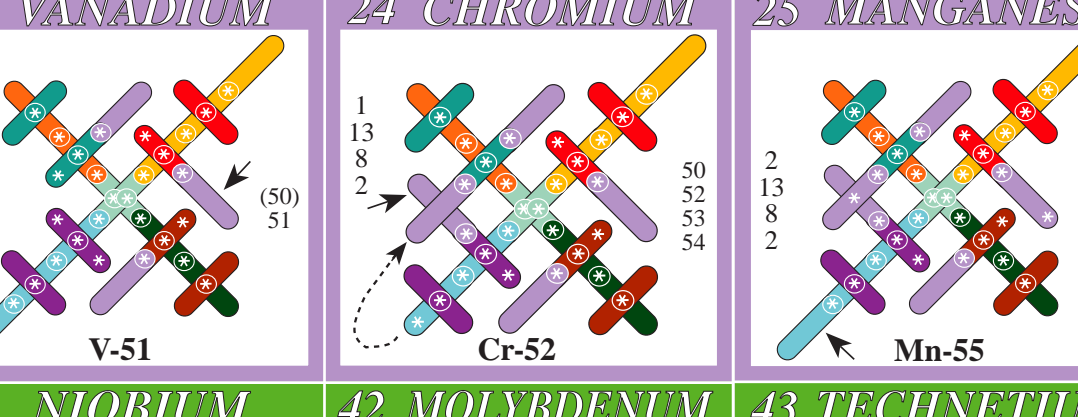
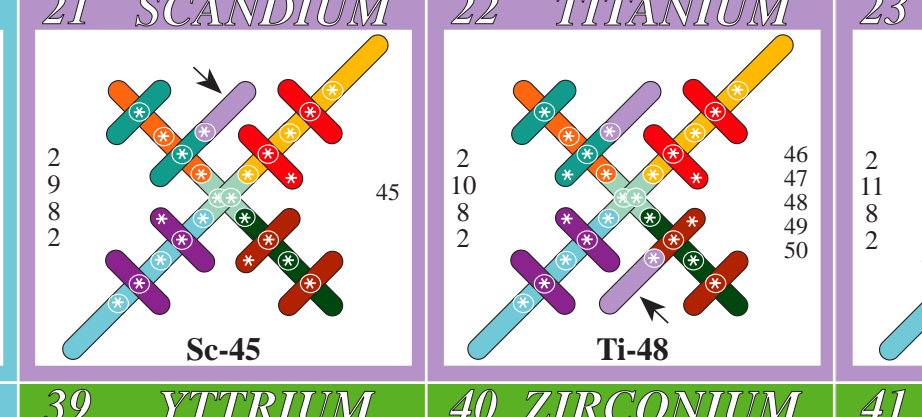
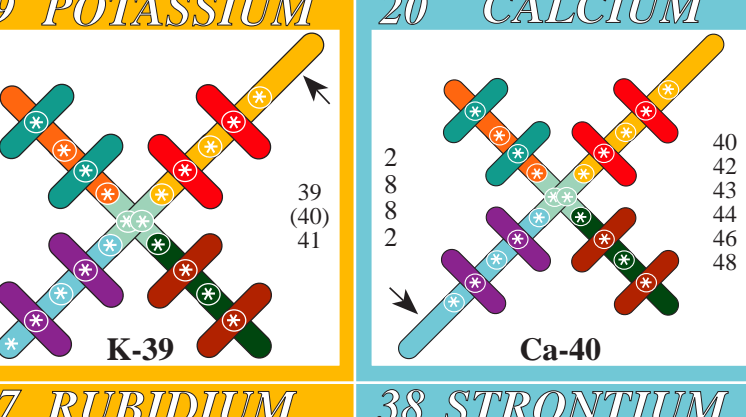
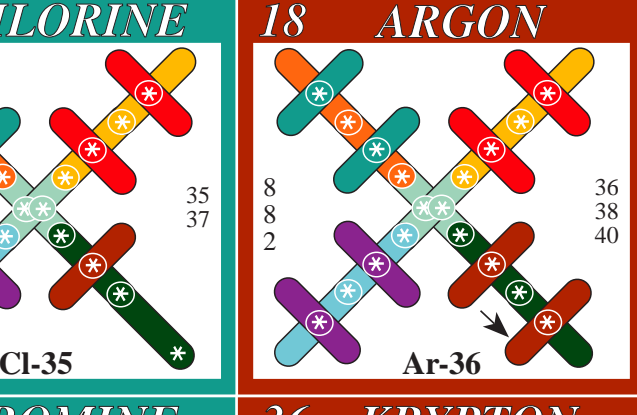
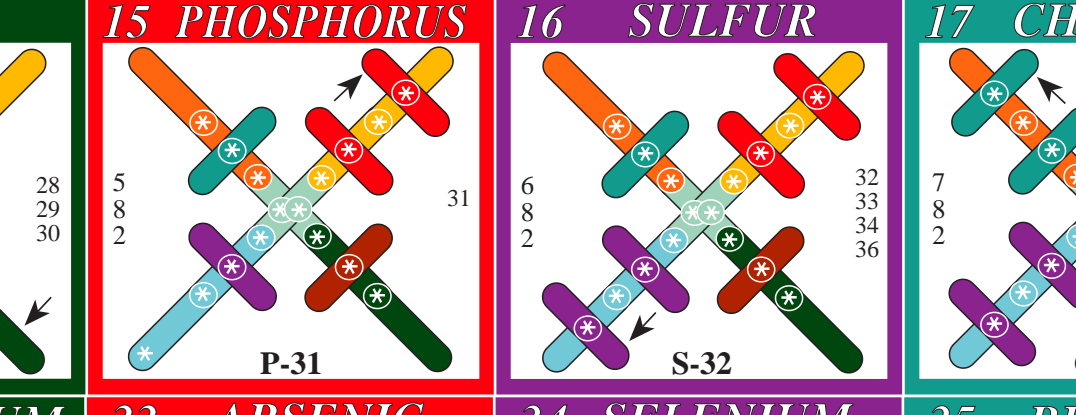
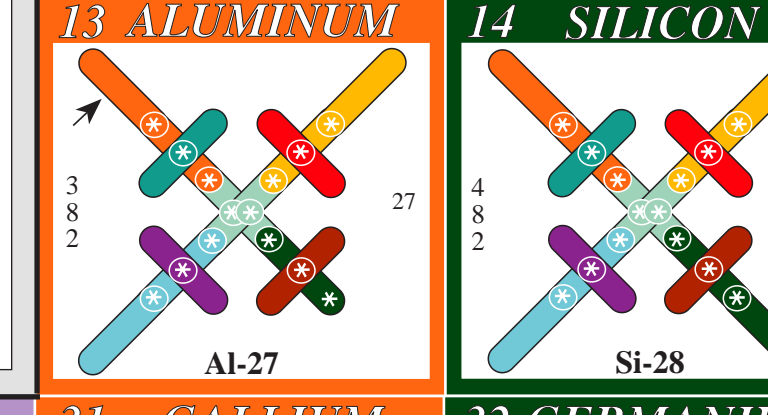
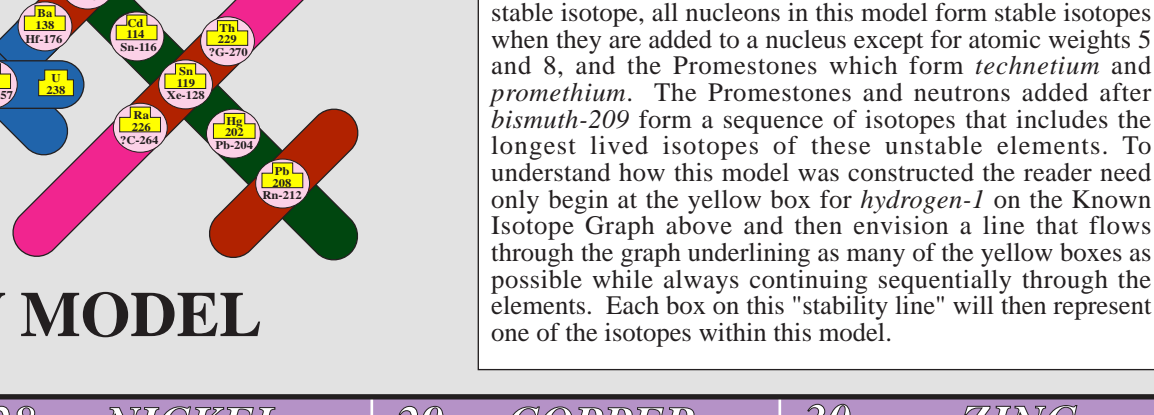
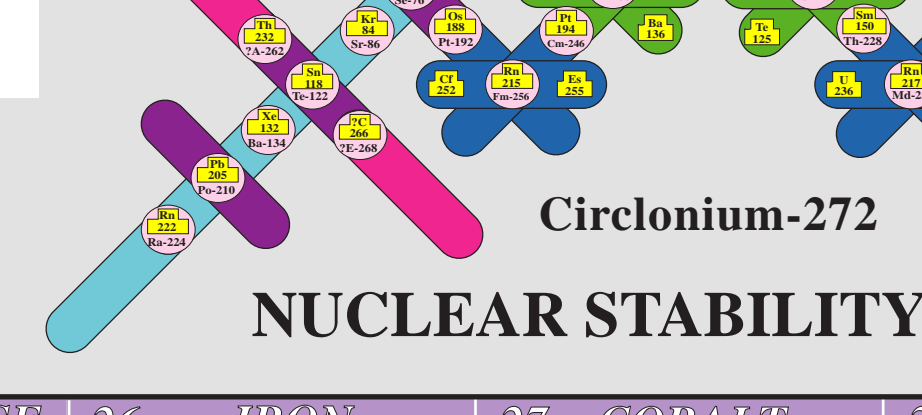
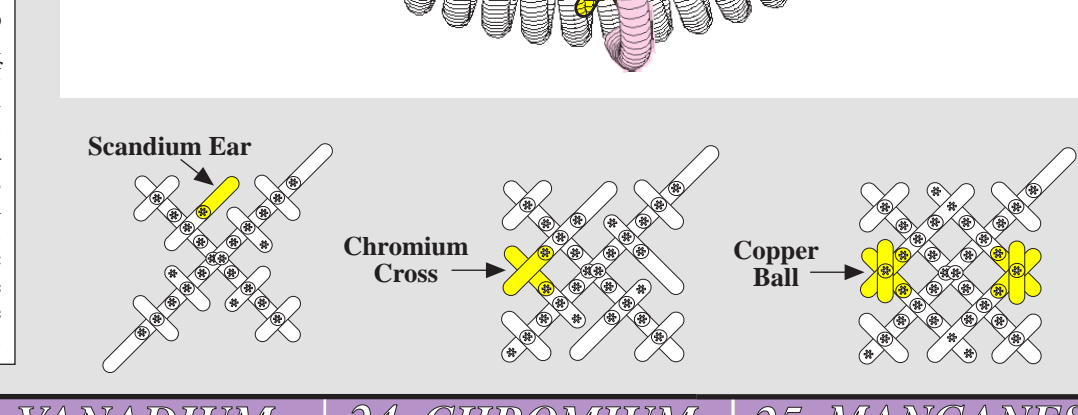
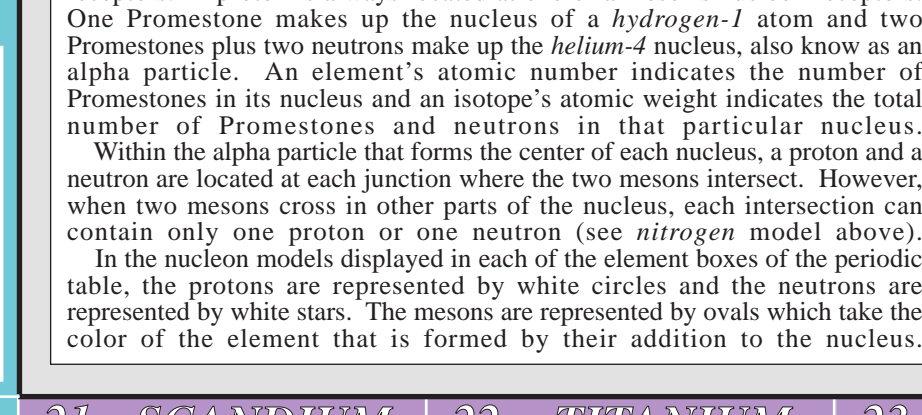
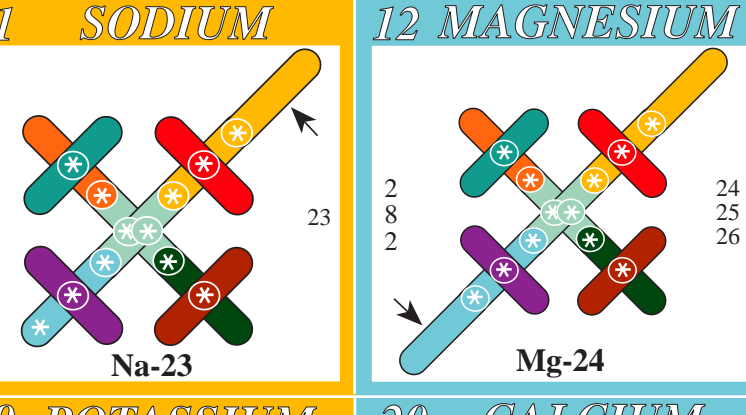
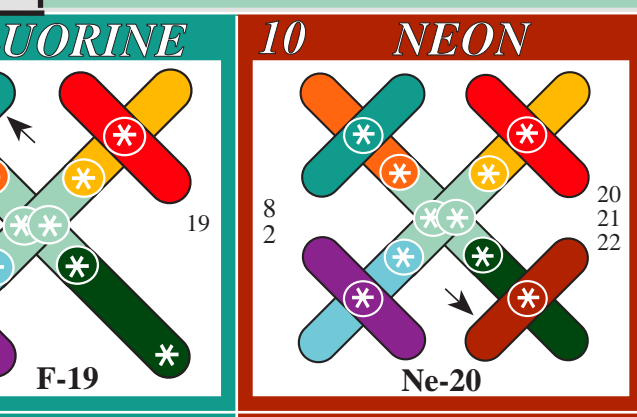
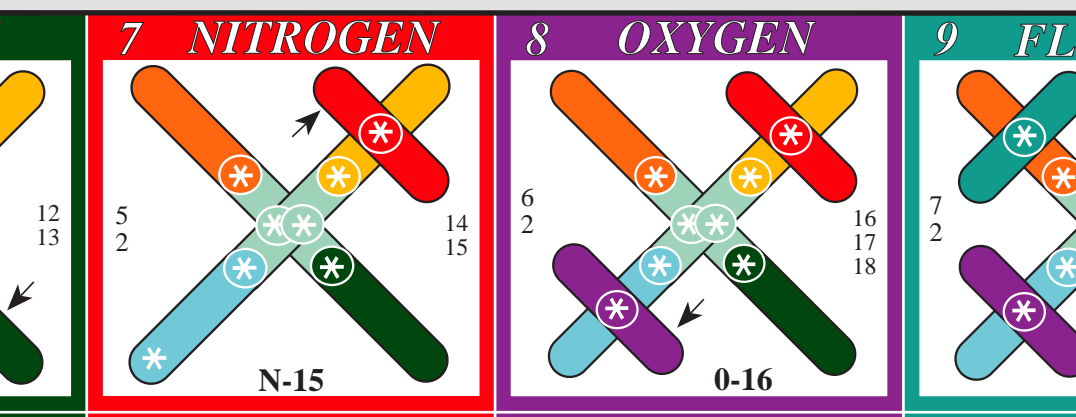
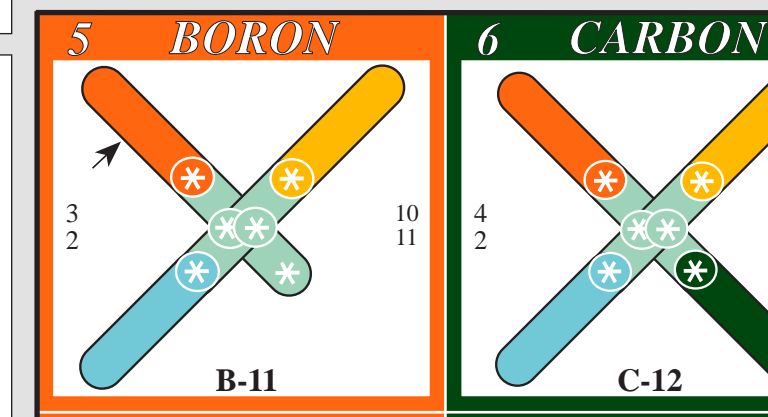
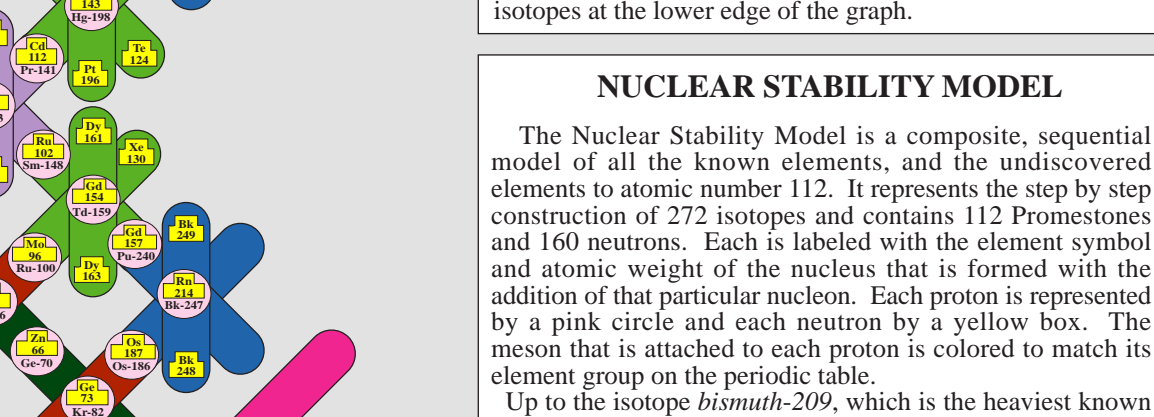
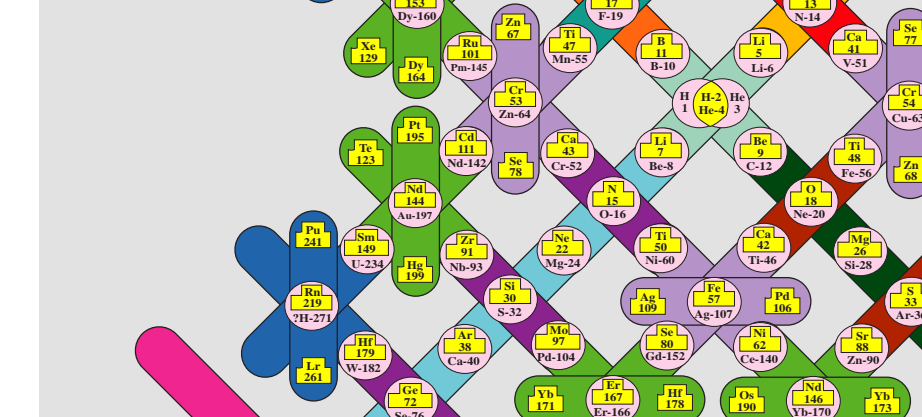
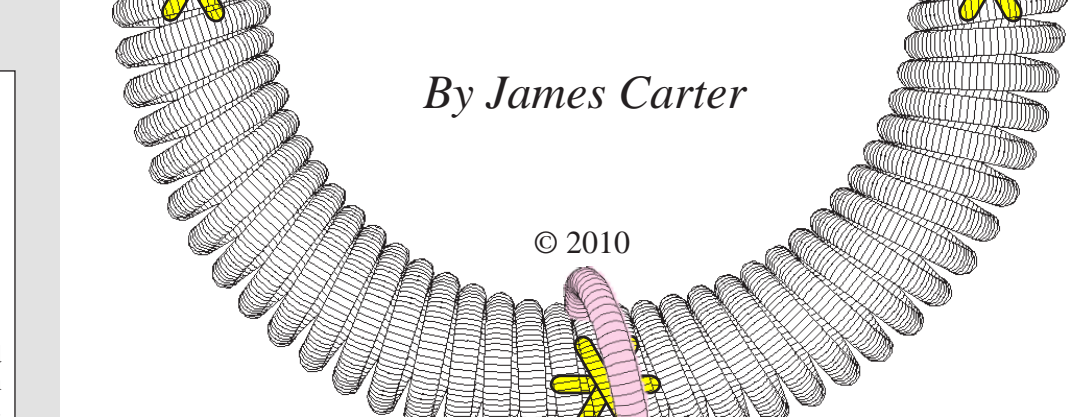
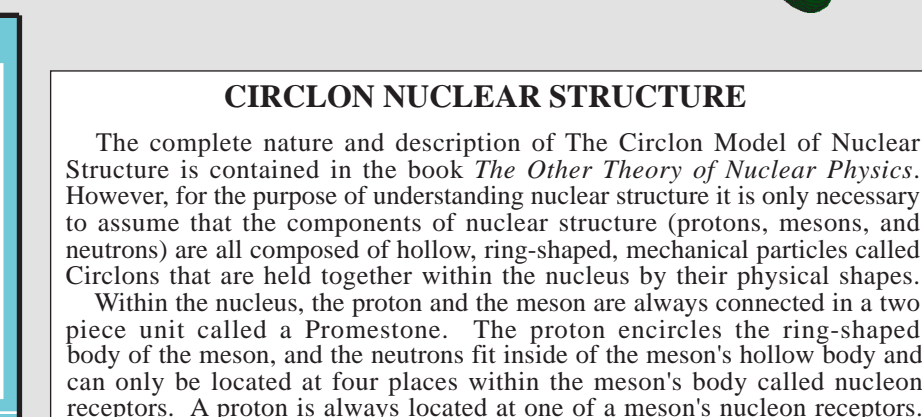
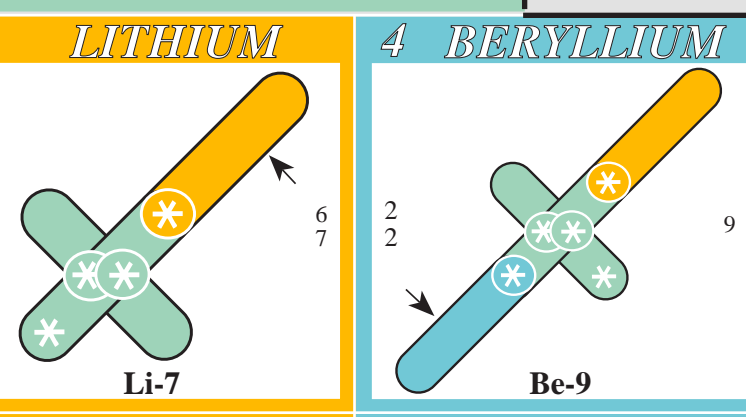
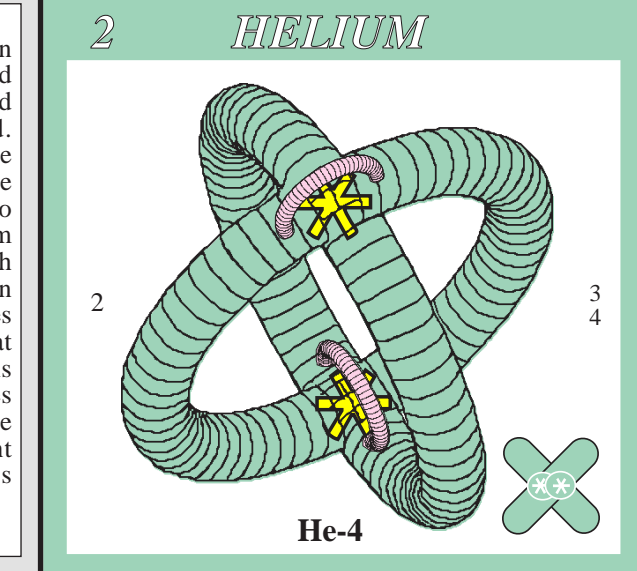


KNOWN ISOTOPE GRAPH

The graph above shows the nearly 2000 isotopes that have been created and identified by researchers. The stable isotopes are identified by yellow boxes and black element symbols. The most common isotope of each element is identified by a yellow line around the yellow box. When an element has no stable isotopes, its longest lived isotope is identified by a yellow line around the isotope box. The row of isotope models at the top of the graph corresponds to the Maximum Structure Line and shows selected elements with their nuclear structure. The yellow neutrons are those which have been added to the Archeopteron of that element (see Maximum Structure Line in *The Other Theory of Nuclear Physics*). The row of isotope models at the bottom of the graph corresponds to the Basic Structure Line and shows selected elements with balance neutrons removed. The yellow neutrons, which appear in those models heavier than rubidium, are balance neutrons that have been added to the basic structure of these nuclei to make them match the isotopes at the lower edge of the graph.

PERIODIC TABLE

The periodic table presented here is divided into 15 different groups, each with its own color. Each vertical row of the Lithium Process element is given a separate color and the elements of each of the four Scandium Processes have different colors. Within each box on the table, a nucleon model of its Archeopteron is displayed for the element. An element's Archeopteron is that isotope which most characterizes the structure of that element and also shares a symmetry with the Archeopteron of the elements around it on the periodic table. In almost all cases, an element's Archeopteron is either its most abundant isotope or its longest lived isotope. The numbers to the right of each element are the atomic weights of the stable isotopes. The numbers to the left of each element (parenthesis indicate radioactivity).



THE OTHER THEORY OF NUCLEAR PHYSICS
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ELECTRON SHELLS

Any new theory, such as the Circlon Model of Nuclear Structure, needs an experiment that can be performed to demonstrate its superiority over the current existing theory. Up until *palladium*, the Circlon Model and the Standard Quantum Mechanical Model of nuclear structure are identical on a quantitative if not a qualitative basis. Beginning with *palladium*, there are discrepancies between the two models in the electron shells predicted by each model. The electron shells, which surround the nucleus, are analogous to its Promestone Layers, since each Promestone holds an electron. The structure of the Promestones within the nucleus should be reflected in the structure of the electron shells surrounding the nucleus. With each of the elements up to *palladium*, the electron shells have been exactly the same for both the Standard Model and the Circlon Model. However, beginning with *palladium*, the electron shells of these two models depart from one another. It is this discrepancy that can make possible an experimental test, which could show one model to be superior over the other. When the transition is made from *ruthidium* to *palladium*, the Standard Model assumes that both the electron from *ruthidium*'s (O) shell and *palladium*'s new electron go down into the (N) shell. However, as we can easily see from the Circlon Model of the nucleus, these two electrons drop down, not into the (N) shell, but rather into the (M) shell. The (M) shell has remained at 18 since the formation of *caesium*. If we look at the model of *Ca-63*, it is easy to see that the first Scandium Process is not complete until its two remaining Chromium (O) shell and *palladium*'s new electron transformed into Copper Balls to complete the Layer of four. These two Copper Balls do not form until *palladium*.

We must reexamine the experiments used to verify the electron configurations as shown in the Standard Model. Without a physical model of the nucleus in question, an experimenter could easily assume that *palladium*'s last two electrons reside in the (N) shell since the (M) shell has been inactive and apparently "full" for 16 consecutive elements. It may be possible, by experimenting with highly ionized *palladium* ions, to determine which of these two electron shell models most accurately describes physical reality. The same situation occurs in the transition from *platinum* to *gold*, so *united gold* could also be used in this experiment.

In the 4 models of elements to the right, the electron shells of the Standard Model are shown in the white boxes and those of the Circlon Model are shown in the colored boxes. The mesons in each nucleus are color-coded to match their respective electron shells.

Na-23

Cu-63

Pd-110

Au-197

